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shown below  $\text{pb} + \text{no}_3^- + 2 \text{aq} + 2 \text{hcl} \text{aq} \rightarrow \text{pbcl}_2 \text{s} + 2 \text{hno}_3$

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web the net ionic equation for equation 1 is given as equation 3  $\text{S}^{2-} \text{ aq} + \text{Fe}^{2+} \text{ aq} \rightarrow \text{FeS} \text{ s}$  eq 3 8 note that the phase indicator aq is optional on ionic species but is

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web  $\text{CuClO}_4$  is soluble all perchlorates are soluble while  $\text{PbSO}_4$  is insoluble sulfates are generally soluble except for  $\text{Pb}^{2+}$   $\text{Ba}^{2+}$   $\text{Hg}_2^{2+}$  and  $\text{Sr}^{2+}$  the net ionic equation is

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